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~~vs unsalted butter Grow~~

~~Purple Single Crystals~~

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*Extracting gold from
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*Copper(II)Sulphate and
Aluminium Reaction*

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Aluminium Cost More
than Gold... *Chemistry*

*Practical Class XII To
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Potash Alum ~~Alum from~~

~~Soda Cans~~ *Make*

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~~(Potassium Aluminium~~

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~~Sulfate~~ *Aluminum and
Mercury Synthesizing
Alum* Exp. 12 -

Synthesis: Preparation
of Alum ~~Alum From
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Alum From Aluminum
Lab Answers The
Synthesis of Alum Lab
Michaela Tonsager and
Kaili Johnson

Conclusion We
determined that our
sample was in fact alum.

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Our melting point of
99.4 degrees C was
similar to the published
melting point of 92.5
degrees C. Our percent
sulfate was 42.44%,
which is close to the
The Synthesis of Alum
Lab by Michaela
Tonsager

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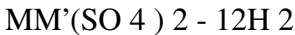
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The aluminum is being recycled in the sense that the aluminum undergoes a process that adapts it to a new function, instead of just converting the shape of the metal. The aluminum in this experiment is converted to alum $[KAl(SO_4)_2 \cdot 12H_2O]$ which is the

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usual term for a type of compound with the general formula



O. M is a monovalent cation, and M' is a trivalent cation, in this case, Al³⁺.

~~Recycling Aluminum~~

~~lab write up: experiment~~

~~3 - CHEM 2070 ...~~

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Lab 4 synthesis and analysis of alum part. Alum from the synthesis of potassium aluminum sulfate chem lab page for the Rinse the reaction beaker twice with 2 mL distilled water, and add each rinse to the filter paper.

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Allow water to run through the aspirator to establish a vacuum in the filter flask. Red numbers are variable.

~~Synthesis of alum
preliminary lab
assignment answers ...~~

1. Either bring your own aluminum can or use the pieces provided in the lab. You will need a piece of scrap aluminum

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about 7.5 cm by 5.0 cm that weighs 1.0 to 1.1 grams. 2. Scrape both sides of the piece (using sandpaper) to remove the paint and lacquer (this will speed the dissolving process GREATLY, and give a more accurate starting mass). 3.

~~Synthesis of Alum from Aluminum~~

Page 16/29

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Synthesis of Alum

Alum is a solid ionic compound with many

uses. It is used as astringent to prevent bleeding from small cuts, as an ingredient in deodorants, as an ingredient in baking powders, and as a preservative used in pickling. The formula of alum is $KAl(SO_4)_2 \cdot 12H_2O$ and the

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standard name is
potassium aluminum
sulfate. It is an

~~CHEM 231 Experiment~~

~~5 Synthesis of Alum~~

Numbers of moles of

Alum = 0.0397 moles

Molar mass of Alum =

474.39 g / mole Mass of

Alum yielded = Number

of moles * Molar mass

of Alum 4. Mass of

Alum yield = 0.0397

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moles * 474.39 (g /
mole) = 18.83 g % yield
= (Actual yield (g) /

Theoretical yield (g)) *
100 = (14.14 g / 18.83
g) * 100 = 75.1 %

Discussion: The
experimental value for
the yielding of alum is
14.14 g and theoretical
value is 18.83 g.

~~Lab report on synthesis
of Alum using~~

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~~Aluminum.~~

In this experiment you will prepare and characterize alum (potassium aluminum sulfate dodecahydrate, $\text{KAl}(\text{SO}_4)_2 \cdot 12 \text{H}_2\text{O}$). The first step in this synthesis, which you will perform during Week 1, is to react metallic aluminum with a concentrated solution of potassium hydroxide

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(KOH) to form the potassium salt of the tetrahydroxoaluminate complex ion, $[\text{Al}(\text{OH})_4]^-$.

~~Preparation and
Analysis of Alum
Chem Lab~~

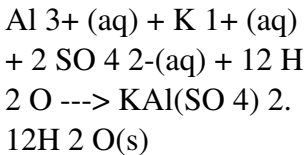
CHEM 231 Experiment
5 Synthesis of Alum
Alum is aluminum sulfate. Aluminum reacts with sulfuric acid

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to produce alum. Chem
Lab: Synthesis of Alum
Question!?! Yahoo

Answers The overall
reaction to form alum is:



Crystallization. Using
tongs, remove the
filtrate beaker to a wire

~~Alum Synthesis Lab~~

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~~Answers~~ Aluminum

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1. Aluminum is the third most abundant element in the earth's crust.

2. The supply of aluminum ore is not inexhaustible. 3. The winning, or extraction of the metallic form an impure ionic source, of aluminum from aluminum ore is very costly from an energy

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point of view. 4. The
above point explains
why Napoleon III used
aluminum dinnerware
for his

~~Synthesis of Alum from
Aluminum~~

stuff is my mid term
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answers the synthesis of
alum lab michaela
tonsager and kaili
johnson conclusion we

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determined that our sample was in fact alum our melting point of 994 degrees c was similar to the published melting point of 925 degrees c our percent sulfate was 4244

~~Analysis Of Alum Lab
Answers Sulfate~~

SYNTHESIS OF

ALUM Purpose: In this lab you are going to

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carry out a series of reactions to transform a piece of aluminum foil into white crystals (alum). You will use stoichiometry to figure out the limiting reagents and yield for your reaction. Background: Aluminum is the third most common element in the earth's crust, but it forms very

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~~SYNTHESIS OF~~

~~ALUM~~

Alum Synthesis Lab

Answers The Synthesis

of Alum Lab Michaela

Tonsager and Kaili

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determined that our
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similar to the published

melting point of 92.5

degrees C. Our percent

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sulfate was 42.44%,

which is close to the

The Synthesis of Alum

Lab by Michaela

Tonsager Lab #4

Synthesis of Alum

Purpose-Synthesize a

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This is a video I made

during my last lab.

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Might get used in a lab report, not sure. The editing in this video is very basic. I just switched from Sony Movie...

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