

## Ap Chemistry Electrochemistry Answers

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**AP Chemistry Electrochemistry Notes** *Electrochemistry Review - Cell Potential \u0026amp; Notation, Redox Half Reactions, Nernst Equation Cell Potential Problems - Electrochemistry*

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AP Chemistry-Electrochemistry. Multiple Choice. Identify the choice that best completes the statement or answers the question. \_\_\_\_ 1. The half-reaction that occurs at the cathode during the electrolysis of molten sodium bromide is \_\_\_\_\_. a.  $+ 2e 2Br \rightarrow Br_2$  b.  $+ 2e Br_2 \rightarrow 2Br^-$  c.  $+ e Na^+ \rightarrow Na$  d.  $Na \rightarrow Na^+ + e^-$  e.  $\rightarrow 2H_2O + 2e^- 2OH^- + H_2$  \_\_\_\_ 2.

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AP Chemistry-Electrochemistry - Quia

AP Chemistry: Electrochemistry Multiple Choice Answers 14. Questions 14-17 The spontaneous reaction that occurs when the cell in the picture operates is as follows:  $2Ag^+ + Cd(s) \rightarrow 2Ag(s) + Cd^{2+}$  (A) Voltage increases. (B) Voltage decreases but remains > zero.

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AP Chemistry: Electrochemistry Multiple Choice Answers

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$\text{Zn(s)} + \text{Ni}^{2+}(\text{aq}) \rightarrow \text{Ni(s)} + \text{Zn}^{2+}(\text{aq})$  (a) Identify M and  $\text{M}^{2+}$  in the diagram and specify the initial concentration for  $\text{M}^{2+}$  in solution.

Electrons flow from the anode to the cathode in a voltaic electrochemical cell. The anode is where oxidation occurs, and in the reaction above,  $\text{Zn(s)}$  is oxidized.

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Electrochemistry - AP Chemistry Advanced Placement Chemistry: 1996

Free Response Questions 7)  $\text{Sr(s)} + \text{Mg}^{2+} \rightleftharpoons \text{Sr} + \text{Mg(s)}$  Consider the reaction represented above that occurs at  $25^\circ\text{C}$ . All reactants and products are in their standard states.

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Electrochemistry - the study of the interchange of chemical and electrical energy There once was a table of reduction potentials in the reference

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AP REVIEW QUESTIONS - Electrochemistry - Answers Answer: (a) tin electrode is the cathode; cathode is the site of reduction (gain in electrons) and will convert metal ions into a metal. (b) (see diagram) (c) red:  $\text{Sn}^{2+}(\text{aq}) + 2 \text{e}^- \rightarrow \text{Sn}(\text{s})$   $E^\circ = -0.14 \text{ V}$  oxid:  $\text{X}(\text{s}) - 3 \text{e}^- \rightarrow \text{X}^{3+}(\text{aq})$   $E^\circ = +0.74 \text{ V}$   $E^\circ \text{ cell} = +0.60 \text{ V}$  red:  $\text{X}^{3+}(\text{aq}) + 3 \text{e}^- \rightarrow \text{X}(\text{s})$

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AP REVIEW QUESTIONS Electrochemistry - Answers

Advanced Placement Chemistry: 1996 Free Response Questions 7)  $\text{Sr}(\text{s}) + \text{Mg}^{2+} \rightleftharpoons \text{Sr} + \text{Mg}(\text{s})$  Consider the reaction represented above that

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occurs at 25°C. All reactants and products are in their standard states. The value of the equilibrium constant,  $K_{eq}$ , for the reaction is  $4.2 \times 10^{17}$  at 25°C.

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A.P. Chemistry Practice Test - Ch. 17: Electrochemistry A ...  
Practice: Electrochemistry questions. This is the currently selected item. Electrochemistry. Redox reaction from dissolving zinc in copper sulfate. Introduction to galvanic/voltaic cells. Electrodes and voltage of Galvanic cell. Shorthand notation for galvanic/voltaic cells.

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Electrochemistry Involves TWO MAIN TYPES Of Electrochemical Cells : 1. Galvanic (voltaic) cells - which are thermodynamically favorable chemical reactions (battery) 2. Electrolytic cells - which are thermodynamically unfavorable and require external e- source (a direct current or DC power source)

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AP\* Chemistry ELECTROCHEMISTRY

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AP Chemistry Review Questions - Electrochemistry. For the galvanic cell described below, the correct line notation is:  $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$  ( $E^\circ = 1.36\text{v}$ )  $\text{Cu} + \text{e}^- \rightarrow \text{Cu}$  ( $E^\circ = 0.52\text{v}$ )  $\text{Cu (s)}|\text{Cu}^+ \text{(aq)}||\text{Cl}_2 \text{(g)}|2\text{Cl}^- \text{(aq)}|\text{Pt (s)}$   $\text{Pt (s)}|\text{Cu (s)}|\text{Cu}^+ \text{(aq)}||\text{Cl}_2 \text{(g)}|2\text{Cl}^- \text{(aq)}|\text{Pt (s)}$

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Answer the following questions regarding the electrochemical cell shown above. (a) Write the balanced net-ionic equation for the spontaneous reaction that occurs as the cell operates, and determine the cell voltage. (b) In which direction do anions flow in the salt bridge as the cell operates? Justify your answer. (c) If 10.0 mL of 3.0-molar AgNO

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