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the SI unit representing 6.02×10^{23} representative particles of a substance: Avogadro's number: 6.02×10^{23} particles: standard temperature and pressure (00 C, 1 atm) the temperature and pressure at which one mole of gas occupies a volume of 22.4 L: molar volume: volume of a gas that contains one mole of the gas, is 22.4 L at STP: Avogadro ...

~~Quia - Chapter 10 "Chemical Quantities"~~

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Vocab. the SI unit representing 6.02×10^{23} representative particles of a substance. the temperature and pressure at which one mole of gas occupies a volume of 22.4 L. equal

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volumes of gases at the same temperature and pressure contain equal numbers of particles.

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Quantities" Vocab~~

Chapter 10 Chemical Quantities 91

SECTION 10.1 THE MOLE: A
MEASUREMENT OF MATTER (pages
287–296) This section defines the
mole and explains how the mole is
used to measure matter. It also
teaches you how to calculate the mass
of a mole of any substance.

~~SECTION 10.1 THE MOLE: A
MEASUREMENT OF MATTER (pages
287–296)~~

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Multiple Choice Identify the choice that best completes the statement or answers the question. _____ 1. What is standard temperature and pressure, or STP? a. 0°C and 101.3 kPa b. 0°C and 1 Pa c. 20°C and 1 Pa d. 100°C and 101.3 kPa _____ 2. At STP, a student measures the volume of 1.00 mole of a gas to be 22.4 L . Based on this measurement, what can the

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The Mole. Study Guide and

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Section 10.3 Chemical Formulas.

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