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answer for each question is indicated by a .
1: Hyperventilation can cause ...

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454 CHAPTER 15 Formula Acid name HF

hydrofluoric acid HCl hydrochloric acid HBr

hydrobromic acid HI hydriodic acid H₂S

hydrosulfuric acid TABLE 15-1 Names of Binary Acids FIGURE 15-2 A strip of pH paper dipped into vinegar turns red, showing that vinegar is an acid.

CHAPTER 15 Acids and Bases

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set (43) All of the following lipids can be hydrolyzed to form fatty acids except ? Steroids. What is a saturated fatty acid? They do not contain a C=C double bond in the carbon chain. What characterizes a triacylglycerol?

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15: Acids and Bases React Interactive Tutor.
True and False Game

Interactive Tutor

Chapter Fifteen: Acid-Base Titration and pH
If you look around the room while titrations

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are happening, there are people standing like this: and the best part is that you probably will too, because it works.

Chapter Fifteen [Acid-Base Titration and pH]

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Chapter 14/15 Equilibrium and Acid-Base

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Practice Test. Multiple Choice (Choose the best answer.--Answers Found At End of Test)

_____ 1. The reaction: $3 \text{CuCl}_2 (\text{aq}) + 2 \text{Al} (\text{s}) \rightleftharpoons 3 \text{Cu} (\text{s}) + 2 \text{AlCl}_3 (\text{aq})$ has the following K_c expression: . a) $\frac{[\text{CuCl}_2]^3 [\text{Al}]^2}{\{ [\text{Cu}]^3 [\text{AlCl}_3]^2 \}}$ b) $\frac{[\text{Cu}]^3 [\text{AlCl}_3]^2}{\{ [\text{CuCl}_2]^3 [\text{Al}]^2 \}}$ c) $\frac{[\text{CuCl}_2]^3}{[\text{Cu}]^3}$ d) $\frac{[\text{Al}]^2}{[\text{AlCl}_3]^2}$ e) None of these

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1. Aqueous solutions of acids have a sour taste. 2. Acids change the color of acid-base indicators. 3. Some acids react w/active metals to release hydrogen gas, H_2 . 4. Acids react with bases to produce salts and water. (salt- the reaction products are water and an ionic compound called a salt). 5. Some acids conduct electric current. 6 ...

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1. strong acid + strong base: steepest at equivalence, low pH at start depending on concentration of acid, equivalence at pH 7, no buffering. weak acid + strong base: less steep at equivalence than strong acid, low pH at start depending on and concentration of acid, equivalence at basic pH, buffering. weak base + strong acid: less steep at equivalence than strong base, high pH at

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start depending on and concentration of base, equivalence at acid pH, ...

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