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Heat of Reaction |  
Enthalpy Of Formation  
| Enthalpy Of  
Combustion | 90

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As liquids absorb heat at  
their boiling points, the

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temperature remains  
constant while they  
vaporize. true \_\_\_\_\_

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Questions 5, 6, and 7.

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g of  $H_2$  and an excess of

$O_2$  are compressed into

a bomb calorimeter

containing 1200 g of

water. The temperature

before the reaction is

$25.00^\circ C$ , and after the

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reaction it goes to 27.16  
°C.  
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Heating Curve for  
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~:;>,-----,1 ~ Melting E

point ~ <-- -----'l Heat

supplied 5. Label the

melting point and

boiling point

temperatures on the

graph. 6. What happens

to the temperature

during melting and

vaporization?

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10. Complete the enthalpy diagram for the combustion of natural gas. Use the thermochemical equation in the first paragraph on page 517 as a guide. SECTION 17.3 HEAT IN

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CHANGES OF

STATE (pages

520 – 526) This section

explains heat transfers

that occur during

melting, freezing,

boiling, and condensing.

SECTION 17.1 THE

FLOW OF ENERGY

HEAT AND WORK

(pages 505 – 510)

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1. How much energy  
must be absorbed by  
20.0 g of water to  
increase its temperature  
from  $283.0\text{ }^{\circ}\text{C}$  to



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303.0 ° C? 2. When  
15.0 g of steam drops in  
temperature from 275.0  
° C to 250.0 ° C, how  
much heat energy is  
released?

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