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Revision Part 2: Determination of the

Molarity of a Sulfuric Acid Solution

strans
Trial 1 Trial 2 Trial 3 Trial 4* Starting Buret

Reading: | • 74 L On15L 0-59L Ending

Buret Reading: 11.83 L 10-10L 10.591 mL of

Titrant Used: 10.09L/9.95L 10.00L mmole of

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NaOH: 1.009 mol / 9.95 g / 1.0 ml / mmole of
H₂SO₄ 0.5 mol / 5.0 g / 0.5 ml / 5.0 ml
H₂SO₄ Solution 0.05 M / 0.5 M / 0.05 M
0.05 M Average Molarity 0.6 M.

Answered: 103 Chemistry I Laboratory
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Grades: The lab component counts as 25%

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of your final grade in CHEM 103. Per department policy, you must perform every experiment in order to pass CHEM 103.

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Reaction 1: $1 \text{ NaHCO}_3 (\text{s}) + 1 \text{ NaOH} (\text{s}) + 1$
 $\text{CO}_2 (\text{g})$ Reaction 2: $2 \text{ NaHCO}_3 (\text{s}) + 1 \text{ Na}_2$
 $\text{CO}_3 (\text{s}) + 1 \text{ CO}_2 (\text{g}) + 1 \text{ H}_2\text{O} (\text{g})$

Reaction 3: $2 \text{ NaHCO}_3 (\text{s}) + 1 \text{ Na}_2\text{O} (\text{s}) +$

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2. Propose a voltaic cell that would coat a piece of copper with platinum. Include the half reactions for your cell and the voltage potential based on the.

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can be purchased (Wiscard only) in room 1375 (chemistry computer lab) during the first two weeks of classes and later through the first-floor stockroom. Lab notebooks (with carbonless copies) purchased from the University Bookstore are also acceptable.

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111. Band-aids for minor ... The answer is
your local wastewater treatment facility,
which operates 24/7 to make sure

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1. Identify the group for each unknown.

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Your unknowns will be from three different groups. (Example: A K^+ salt, a halogen, and an acid.) 2. Identify the specific compound from each group. (Example: KCl , Br_2 and HCl) 3. Explain the reasoning that enabled you to come to these conclusions.

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