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organic chemistry is the study of carbon compounds. Each section of the chapter addresses a different group of organic compounds, related by structure.

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Each section also features important reactions of each compound ...

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Chapter 3 Quiz.

True/False. Indicate whether the sentence or statement is true or false. 1. The region in space where an electron is most likely to be found is called an energy level. 2.

Electron

configurations are often condensed by

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writing them using
the previous noble
gas core as a starting
point. ...

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laboratory section in which we will explore micro and macro levels of chemistry.

We will explore exciting topics as atomic structure, the periodic table ...

~~Welcome to Dr.
Mintz's Chemistry
Classroom - Home
Solution: OH OH OH
0.50 mmol mL - -~~

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$n \times V = 12.5$

mL OH⁻ = 12.5

mmol HCl(aq) HCl(aq)

[HCl(aq)] 12.5 mmol

= $n \times V$ 0.10 mmol

HCl(aq) /mL $V = 120$

mL 2. (a) Given: V

HCl(aq) = 25.00 mL;

[HCl(aq)] = 0.350

mol/L; V NaOH(aq) =

5.0 mL; [NaOH(aq)] =

0.500 mol/L;

Required: [H⁺(aq)],

pH Analysis: Before

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titration, amount of
H⁺ is: H(aq) H(aq)
H(aq) 0.350 mmol mL
++ += x = nc V

~~Section 8.7:~~

~~Acid Base Titration~~

~~Tutorial 1 Practice,~~

~~page 547~~

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Identify any errors in the structure by drawing them.

Rename the structure correctly. 2-ethyl-2,4,4-trimethylpentane

What Is Required?

You are to identify the errors in the name by drawing the structure and

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renaming it correctly.

What Is Given? You are given an incorrectly named alkane.

~~Unit 1 Organic~~

~~Chemistry Mr.~~

~~Arthur's Science Page~~

~~= 0.12 mol/L~~

~~Statement: The~~

~~equilibrium~~

~~concentration of~~

~~dinitrogen tetroxide~~

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gas is 0.12 mol/L. 2.

Given: initial quantity
of $\text{NOCl(g)} = 3.00$

mol; volume = 3.00 L;

$[\text{NO(g)}]$ initial = 0.00

mol/L; $[\text{Cl}_2(\text{g})]$ initial

= 0.00 mol/L; $[\text{NO(g)}]$

equilibrium = 0.043

mol/L Required:

$[\text{NOCl(g)}]$

equilibrium; $[\text{Cl}_2(\text{g})]$

equilibrium Solution:

Step 1.

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~~Section 7.1:~~

~~Equilibrium Systems~~

~~Mini Investigation:~~

~~The ...~~

Solution: bonds

broken 3 mol 1 mol

CH CCl 3 mol HnD DD

- - = x = x + x =

413 kJ mol x +1 mol

339 kJ mol x H

=1578 kJ Statement:

The energy required

to separate 1 mol of

chloromethane into

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free atoms is 1578 kJ.

2. Solution: Step 1:
Write the balanced
chemical equation for
the combustion of 1
mol of ethene gas. $C_2H_4(g) + 3 O_2(g) \dots$

~~Section 5.3: Bond
Energies Tutorial 1
Practice, page 312~~
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Solution: $q = Ne = (1.2!$

$108)(1.602!10^{19} \text{ C})$

$q = 1.92!10^{11} \text{ C}$ (one

extra digit carried) $F E$

$= kq_1 q_2 / r^2 =$

$8.99!10^9 \text{ N}\cdot\text{m}^2 / \text{C}^2 \cdot$

$\% \& ' ((1.92!10)^{11}$

$\text{C})(1.92!10)^{11} \text{ C} (1.10$

$\text{m})^2 F E = 2.7!10^{12} \text{ N}$

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Statement: The force of repulsion between the two plastic spheres is 2.7×10^{-12} N. 2. Given: $m = 2.48 \times 10^{-15}$ kg; $d = 1.7$ cm = 0.017 m; $V_b = 260$ V; $e = 1.602 \times 10^{-19}$ C; $g = 9.8$ m/s² Required: q ; N

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