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Chemistry I: Chapter 14 Test, Gases  $PV=nRT$   $P_1V_1/T_1 = P_2V_2/T_2$   
760 mm Hg = 1 atm  $K = ^\circ C + 273$  1000 mL = 1 L  
22.4 L = 1 mol (at STP)  $R = 0.0821 \text{ atm}\cdot\text{L}$   $R = 8.314 \text{ kPa}\cdot\text{L}$   
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