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The primary objective of chemical thermodynamics is the establishment of a criterion for determination of the feasibility or spontaneity of a given transformation. In this manner, chemical thermodynamics is typically used to predict the energy exchanges that occur in the following processes: Chemical reactions; Phase changes; The formation of solutions

Chemical thermodynamics - Wikipedia

Given that the free energy of formation of liquid water is -237 kJ / mol , calculate the potential for the formation of hydrogen and oxygen from water. To solve this problem we must first calculate ΔG for the reaction, which is $-2(-237 \text{ kJ / mol}) = 474 \text{ kJ / mol}$. Knowing that $\Delta G = -nFE^{\circ}$ and $n = 4$, we calculate the potential is -1.23 V .

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Chemical thermodynamics Wikipedia. Fundamentals of Thermodynamics.

Problems And Solutions In Thermodynamics

The propensity for any two substances to form a solid solution is a complicated matter involving the chemical and crystallographic properties of the substances as well as temperature and pressure. At higher temperatures substances have greater atomic vibration and more open structures which are easier to distort locally to accommodate differently sized cations.

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