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any stoichiometry problem, the first step is always to calculate the number of moles of each reactant present. In this case, we are given the mass of  $K_2Cr_2O_7$  in 1 mL of solution, which we can use to calculate the number of moles of  $K_2Cr$  ...

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chem calculation review. What is the  
mass in grams of  $15 \times 10^{16}$  atoms s.

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reacts with oxygen to produce  
aluminum oxide as follows:  $4\text{Al} + 3\text{O}_2$   
 $\rightarrow 2\text{Al}_2\text{O}_3$  a. If you use 2.3 moles of



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**Key** a. how many moles of  $\text{Al}_2\text{O}_3$  can you make? b. If you want 3.9 moles of  $\text{Al}_2\text{O}_3$ , how many moles of  $\text{O}_2$  are needed? 2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate:  $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$

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Stoichiometry. that portion of  
chemistry dealing with the numerical  
relationships in chemical reactions .

What is stoichiometry based on? the

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Key law of conservation of mass. What does stoichiometry involve?  
balancing chemical equations and mole ratios. mole ratio. a conversion factor that relates the number of moles of any two ...

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